



# UNITED STATES PATENT AND TRADEMARK OFFICE

UNITED STATES DEPARTMENT OF COMMERCE  
United States Patent and Trademark Office  
Address: COMMISSIONER FOR PATENTS  
P.O. Box 1450  
Alexandria, Virginia 22313-1450  
[www.uspto.gov](http://www.uspto.gov)

APPLICATION NO.	FILING DATE	FIRST NAMED INVENTOR	ATTORNEY DOCKET NO.	CONFIRMATION NO.
09/718,767	11/22/2000	Tsuyonobu Hatazawa	09792909-4673	2706
26263                      7590                      01/21/2009 SONNENSCHN NATH & ROSENTHAL LLP P.O. BOX 061080 WACKER DRIVE STATION, SEARS TOWER CHICAGO, IL 60606-1080				
EXAMINER CREPEAU, JONATHAN				
ART UNIT		PAPER NUMBER		
1795				
MAIL DATE		DELIVERY MODE		
01/21/2009		PAPER		

**Please find below and/or attached an Office communication concerning this application or proceeding.**

The time period for reply, if any, is set in the attached communication.



UNITED STATES PATENT AND TRADEMARK OFFICE

Commissioner for Patents  
United States Patent and Trademark Office  
P.O. Box 1450  
Alexandria, VA 22313-1450  
[www.uspto.gov](http://www.uspto.gov)

**BEFORE THE BOARD OF PATENT APPEALS  
AND INTERFERENCES**

Application Number: 09/718,767  
Filing Date: November 22, 2000  
Appellant(s): HATAZAWA ET AL.

\_\_\_\_\_  
David R. Metzger  
For Appellant

**EXAMINER'S ANSWER**

This is in response to the appeal brief filed November 4, 2008 appealing from the Office action mailed July 1, 2008.

**(1) Real Party in Interest**

A statement identifying by name the real party in interest is contained in the brief.

**(2) Related Appeals and Interferences**

The examiner is not aware of any related appeals, interferences, or judicial proceedings which will directly affect or be directly affected by or have a bearing on the Board's decision in the pending appeal.

**(3) Status of Claims**

The statement of the status of claims contained in the brief is correct.

**(4) Status of Amendments After Final**

The appellant's statement of the status of amendments after final rejection contained in the brief is correct.

**(5) Summary of Claimed Subject Matter**

The summary of claimed subject matter contained in the brief is correct.

**(6) Grounds of Rejection to be Reviewed on Appeal**

The appellant's statement of the grounds of rejection to be reviewed on appeal is correct.

**(7) Claims Appendix**

A substantially correct copy of appealed claim 1 appears on page 15 of the Appendix to the appellant's brief. The minor errors are as follows: in the third clause from the bottom of page 15, the claim does not reflect the amendment made on April 1, 2008 and should read, "said first and second gas absorbable members each being a continuous solid plate-like member and each comprising a gas absorbable material powder hardened within a resin material;"

**(8) Evidence Relied Upon**

EP 895296	NAKANE	2-1999
5,445,856	CHALONER-GILL	8-1995
5,219,676	BULLOCK ET AL.	6-1993
5,607,485	GOZDZ ET AL.	3-1997
4,269,905	WEDLAKE	5-1981

**(9) Grounds of Rejection**

The following ground(s) of rejection are applicable to the appealed claims:

Claims 1, 5, and 8-11 are rejected under 35 U.S.C. 103(a) as being unpatentable over EP 895296 in view of Chaloner-Gill (U.S. Patent 5,445,856) in view of Bullock et al (U.S. Patent 5,219,676) in view of Gozdz et al (U.S. Patent 5,607,485).

Regarding claim 1, EP '296 teaches a nonaqueous electrolyte battery comprising a battery element (6) is contained in an outer covering member including a laminated film having an outermost layer and is sealed by heat seals. The laminate may be in the form of a single sheet, with the first and second covering members folded together and heat-sealed. The battery comprises an inorganic fine oxide powder such as silica which may be located between the battery element and the casing (see [0031], [0033]). It is disclosed that the inorganic oxide functions to "effectively absorb the hydrofluoric acid and others." However, it is submitted that the oxide powder, in aggregate form, would also inherently function as a "gas absorbable

member” as recited in claim 1. The battery electrodes contain carbon and a lithium composite oxide (see [0027]).

The reference does not expressly teach that the oxide powder is a continuous solid plate-like member forming first and second gas absorbable members, as recited in claims 1 and 5.

However, it is submitted that the artisan would be motivated to mix the inorganic oxide powder of EP ‘296 with a binder material to form solid plate-like continuous gas absorbable members. In [0031], the reference teaches that the powder may be present in the space between the battery in the case, but may also be present as an electrolyte or electrode additive. The artisan would recognize from this disclosure that it would be advantageous to include a binder such as resin for the loose powder present in the space between the case and the battery. Such a resin binder would allow the gas absorbable members to become self-supporting and would prevent migration of the powder to other parts of the battery. Accordingly, the use of resin to form gas absorbable members would be obvious to the skilled artisan.

EP ‘296 further does not expressly teach a winding type battery element as recited in claim 1, or that the outer laminated covering member contains a gas absorbable material comprising a molecular sieve or a silica gel mixed with a resin material, as also recited in claim 1.

Chaloner-Gill teaches a spirally-wound “jelly roll” type cell (see col. 3, line 45). As shown in Figure 1, a battery element (10) is contained in an outer covering member including a laminated film having an outermost layer and is sealed by heat seals. The outer covering member may be defined as the laminate including layers 64, 66, 68, 70, and 72 (see Fig. 5), and each outer covering member has a recess therein (see Figs. 3 and 4). The battery comprises a gas

absorbing material which is mixed with a resin material and formed as a layer (62) between the “outer covering member” as defined above and the battery element (see col. 2, line 61; col. 6, line 33). In the preferred embodiments, combined thickness of the seven layers (60-72) is less than about 500 microns or less than about 250 microns (see col. 9, line 39). In the case of the latter, the thickness of each layer is preferably 20-30 microns (see col. 9, line 42). The gas absorbing material may comprise an activated carbon material (see col. 8, line 13) or silica gel (see col. 9, line 11).

It is submitted that the artisan would be motivated to use the laminated outer covering member of Chaloner-Gill, which includes the silica gel gas absorbable material, in the battery of EP ‘296. In column 1, line 49, Chaloner-Gill discloses that “[a] laminate is provided for use as a protective covering for inhibiting penetration of oxygen and/or oxygen and water therethrough and is particularly suitable for protecting components of an electrochemical cell such as a lithium battery.” Accordingly, the artisan would be motivated to use the laminated outer covering member of Chaloner-Gill, which includes the silica gel gas absorbable material, in the battery of EP ‘296.

Further, the use of a spirally-wound battery element as disclosed by Chaloner-Gill in the battery of EP ‘296 would be obvious to the skilled artisan. It is known that the use of a wound configuration allows the current density of the battery to be increased. As such, the use of a wound battery element in the battery of EP ‘296 would be rendered obvious.

Chaloner-Gill does not expressly teach that the gas absorbable material is present in an amount of 0.1 to 95 wt. percent on a basis of a weight of the resin material, or that the layer

containing the gas absorbable material has a thickness of between 1-500 microns, as recited in claim 1.

However, the artisan would be motivated to use a suitable amount of gas absorbing material based on the size of the battery and/or electrode element, thereby rendering the claimed range obvious. It has been held that the discovery of an optimum value of a result effective variable in a known process is ordinarily within the skill of the art. *In re Boesch*, 205 USPQ 215 (CCPA 1980). In this case, it is known that an amount of gas absorbing material can be selected based upon the size of the battery, as shown by column 6, lines 26-35 of Bullock et al.:

The amount of the gel will vary from battery size to battery size within the guidelines of being a sufficient quantity to absorb all water vapor produced during the self-discharge reactions. For normal 12V automobile batteries having six cells, the amount could range between about 50 grams to 300 grams. One skilled in the art could readily select a desiccant quantity by knowing the battery size, plate construction and volume of electrolyte left in the battery after the dumping step.

Although Bullock relates to lead-acid batteries, its teachings regarding the battery size would be applicable to all batteries employing a gas absorbing agent. Accordingly, the artisan would be motivated to use a suitable amount of gas absorbing material in the battery of Chaloner-Gill, thereby rendering the range of claim 1 obvious.

Additionally, the thickness range recited in claim 1 is not considered to distinguish over the references. Claim 1 recites that the members have a thickness of between 1-500 microns. As noted above, Chaloner-Gill teaches that in a seven-layer construction, each sheet preferably has a thickness of less than about 500 microns (or less than about 250 microns). In the case of 500 microns, each sheet would have a thickness of about 70 microns. As such, this disclosure fairly suggests the claimed range of 1-500 microns.

EP '296 further does not expressly teach that the electrolyte is a gel electrolyte comprising vinylidene fluoride-hexafluoropropylene (PVDF:HFP) copolymer as recited in claim 1.

The patent of Gozdz et al is directed to a lithium secondary battery. The battery may contain a gel electrolyte containing a PVDF:HFP copolymer and 20-70 wt% of a plasticizer containing an electrolytic salt (see abstract).

Therefore, the invention as a whole would have been obvious to one of ordinary skill in the art at the time the invention was made because the disclosure of Gozdz et al. provides the artisan sufficient motivation to use these materials in the battery of EP '296. In column 2, line 30, Gozdz et al. teach the following:

The present invention provides a means for avoiding the disadvantages of prior electrolytic cell compositions and constructions by enabling the ready and economical preparation of strong, flexible polymeric electrolytic cell membranes which will readily retain electrolyte salt solutions and remain functional over a range extending well below room temperature.

As such, the artisan would be motivated to use the electrolyte of Gozdz in the battery of EP '296.

Claim 4 is rejected under 35 U.S.C. 103(a) as being unpatentable over EP '296 in view of Chaloner Gill in view of Bullock in view of Gozdz et al. as applied to claims 1, 5, and 8-11 above, and further in view of Wedlake (U.S. Patent 4,269,905).



Neither EP '296 nor Chaloner-Gill expressly discloses that the battery contains a carbon molecular sieve, as recited in claim 4.

Wedlake is directed to electrochemical cells having a casing containing a layer of molecular sieve material. The molecular sieve may comprise materials such as zeolite and carbon (col. 3, line 59 et seq.).

Therefore, the invention as a whole would have been obvious to one of ordinary skill in the art at the time the invention was made because the artisan would be motivated to use the carbon molecular sieve of Wedlake as the gas absorbable material of EP '296/Chaloner-Gill. In the abstract, Wedlake discloses that "this invention involves associating a micromolecular sieve carrier with the cell to sorb such contents when they escape, to reduce the severity of undesired reactions of such contents." As such, the artisan would be motivated to use the materials disclosed by Wedlake, such as carbon molecular sieve, as the gas absorbable material of EP '296/Chaloner-Gill in hopes of obtaining these advantages.

#### **(10) Response to Argument**

Appellants state that "the *Nakane* powder functions to absorb liquids, namely water and hydrofluoric acid, and not gas, as required by the claims" (Appeal Brief, p. 6). Although Nakane (EP '296) does teach liquid absorption, it is submitted that the powder is at least capable of absorbing a gas, thereby rendering it a "gas absorbable material." First, if the material can absorb water as disclosed by the reference, then it would also be capable of absorbing water in a gaseous form, i.e., steam. Second, the reference expressly discloses that the material may be a

zeolite (see [0033]), which is a species of a molecular sieve. Such molecular sieves are inherently gas absorbable materials, as evidenced by their inclusion in present claim 1 (see third clause of claim 1). Nakane also teaches materials such as alumina, silica, magnesia and titania, which are also inherently capable of absorbing gases (see instant specification, page 6, last paragraph, which expressly discloses that zeolite, alumina, molecular sieve, titania, silica gel and zirconia are examples of gas absorbable materials).

Appellants further state that “*Nakane*’s fine powder is made up of individual particles which do not form a **continuous solid member**” (Brief, p. 6). However, this subject matter is believed to be obvious for the reasons stated in the above rejection. In addition, in the final rejection of July 2008, the Examiner made the following argument:

Further, Nakane, as well as Chaloner-Gill, teach that the powder, when used in other locations in the battery, is combined with a resin (i.e., electrolyte or packaging layer). Therefore, it would be well within the skill of the art to employ a resin to contain the powder that is located between the battery element and package of Nakane. Further, the subject matter would have been obvious because a particular known technique (combining gas absorbable powder with a resin) was recognized as part of the ordinary capabilities of one skilled in the art. *KSR v. Teleflex*, 82 USPQ2d 1385, 127 S. Ct. 1727 (2007). The use of resin to form the gas absorbable material of Nakane into a “plate-like member” is within the skill of the art and would yield a predictable result. Accordingly, the rejection as stated above is believed to be proper.

It should be noted that Appellants have not offered a substantive rebuttal to the Examiner’s stated rationale for employing a resin binder in Nakane. Accordingly, it is respectfully requested that the Board sustain the rejection on these grounds.

Further, in the response of April 1, 2008, Appellants argued that “[a]s described in Applicant’s specification, Applicants’ plate-like members beneficially provide improved volume energy efficiency compared to encasing a battery element with a gas absorbable material. [...] Thus, Applicants’ plate-like members provide improved volume energy efficiency compared to *Nakane*.” However, Nakane does not teach any specific configuration of the powder surrounding

the battery. Thus, Appellant's comparison of the plate-like members to an encasing member is not germane to Nakane. Further, the artisan would be sufficiently skilled so as to not use so much of the gas absorbable material so as to compromise the volumetric efficiency of the battery.

On page 8 of the Brief, Appellants further argue that "[a]s the present application discloses, the gas absorption members of the claimed invention are effective to absorb carbon dioxide, propane, propylene and evaporated propylene carbonate gases which are produced during the operation of the battery and which may collect at the first and second ends of the battery." This argument is noted; however, the claims on appeal do not specify which gas(es) the gas absorbable material must have the capability of absorbing. Further, as stated above, it is submitted that the materials of Nakane are inherently capable of absorbing gas.

Regarding Chaloner-Gill, Appellants argue that the reference "discloses using a laminated layer which **has low permeability** which would **restrict** gas from interacting with the gas absorption material in the layer thereby **inhibiting** the absorption of gas by the gas absorption member" (Brief, p. 9). In response, it is submitted that all of the structural aspects of the claimed configuration have been met by the references. Chaloner-Gill is relied on for its teachings of a battery packaging laminate containing a layer of gas absorbing material in one of the laminate layers. Appellant's argument is directed to the functionality of the laminate, however this does not negate the fact that the claimed structural aspects of the laminate are disclosed by Chaloner-Gill. It is noted that the reference expressly discloses that the laminate comprises a layer having an "oxygen scavenger" mixed with the polymer of the layer (col. 8, line 50).

Regarding the Wedlake reference, Appellants state that "because the Wedlake sieve is directed to absorbing liquids, it is not suitable for absorbing gases as called for in the claims" (Brief, p. 12). However, the carbon molecular sieve disclosed by Wedlake is inherently capable of absorbing gases, as evidenced by its recitation in present claim 4 on appeal. Further, Appellant has not provided any technical reasoning to support the position that a carbon molecular sieve is not capable of absorbing a gas of any composition. In addition, the Examiner's arguments with respect to Nakane above are also applicable to Wedlake.

**(11) Related Proceeding(s) Appendix**

No decision rendered by a court or the Board is identified by the examiner in the Related Appeals and Interferences section of this examiner's answer.

For the above reasons, it is believed that the rejections should be sustained.

Respectfully submitted,

/Jonathan Crepeau/  
Primary Examiner, Art Unit 1795

Conferees:

/PATRICK RYAN/  
Supervisory Patent Examiner, Art Unit 1795

/Dah-Wei Yuan/  
Supervisory Patent Examiner, Art Unit 1795